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$1/1.41 = 0.71 \text{ m}^3/\text{kg}$
 $2.233 \text{ kT} = (1.38 \times 10^{-23}) (500) = 1.035 \times 10^{-20} \text{ J}$
One kg-mole, which has a mass of 4 kg, has 6.02×10^{26} atoms.

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Charles Lindbergh Chair at the National Air and Space Museum of the Smithsonian
Institution. He continued with the Air and Space Museum one day each week as their Special
Assistant for Aerodynamics, doing research and writing on the History of Aerodynamics.

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